

Name Key

WS 4a

Dissociation of Strong and Weak Electrolytes

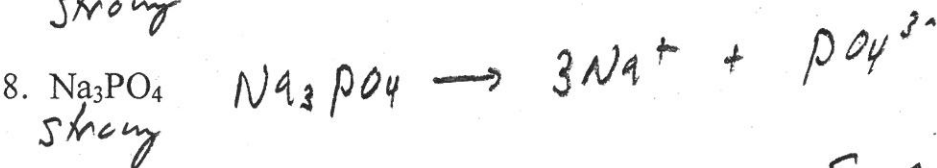
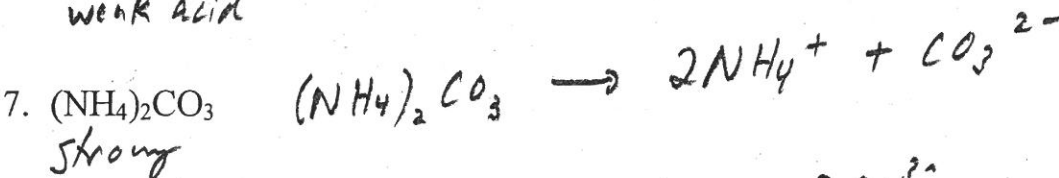
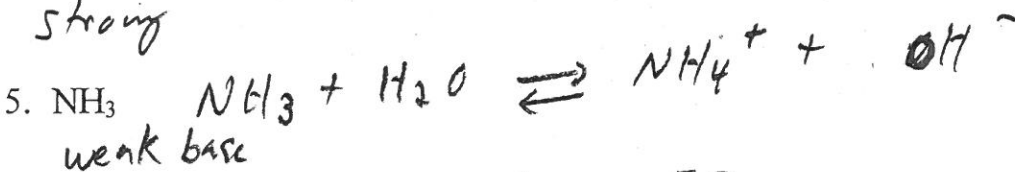
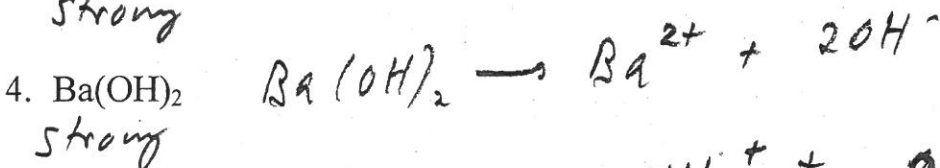
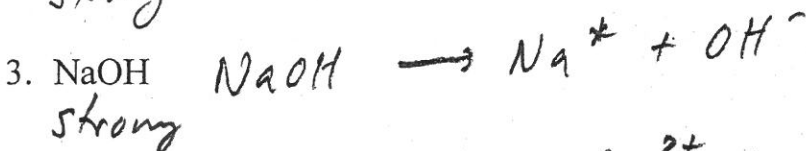
Instructions: Identify each of the substances below as either strong or weak electrolytes. In each case, write a balanced equation that represents how and if the substance dissociates in water. It is NOT necessary to write in the water.

Background: Rules

Strong electrolytes = soluble salts, strong acids and strong bases. Seven common strong acids encountered most often are HCl, HBr, HI, HNO₃, HClO₃, HClO₄, and H₂SO₄. The common strong bases include the Group 1 hydroxides (albeit FrOH), and the heavier group 2 hydroxides, Ca, Sr, and Ba.

Weak electrolytes are most acids (other than those mentioned above) and most bases (other than those mentioned above).

Solubility rules on other salts to memorize: all salts of Na⁺, K⁺, NO₃⁻, and NH₄⁺.



Thiosulfate ion
Strong

