

1. Each of the following pairs of aqueous solutions are mixed.

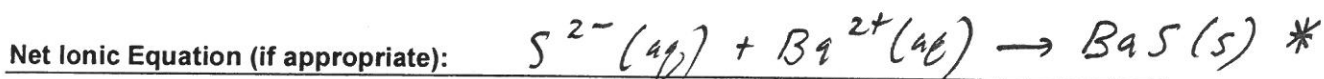
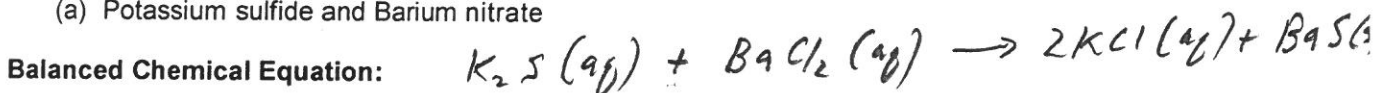
Either

- Write a full, balanced chemical equation for the double displacement reaction that takes place indicating the precipitate formed by adding the (s) state symbol in the equation, and using (aq) state symbols where appropriate, **AND** write the **Net Ionic Equation** including the state symbols.

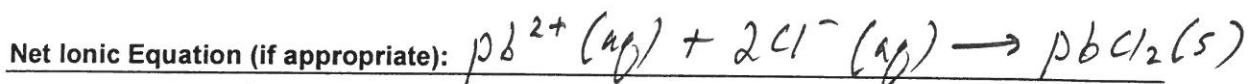
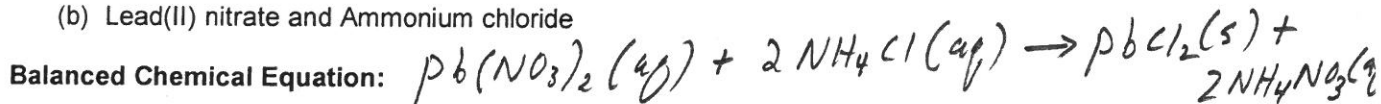
or

- If NO precipitate (solid) forms, write **NO REACTION** instead of a full, balanced chemical equation, **AND DO NOT** write a Net Ionic Equation.

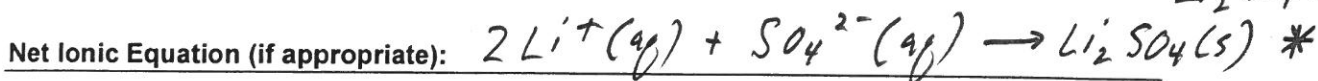
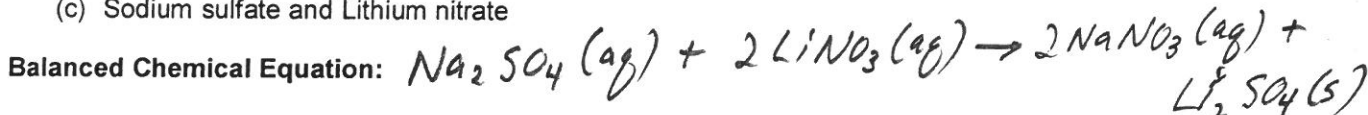
(a) Potassium sulfide and Barium nitrate



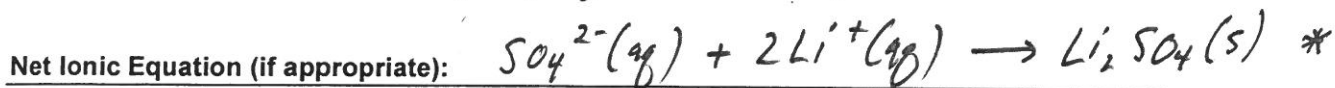
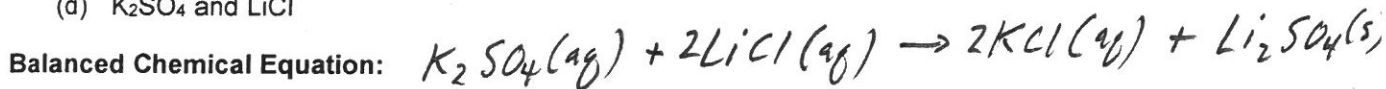
(b) Lead(II) nitrate and Ammonium chloride



(c) Sodium sulfate and Lithium nitrate



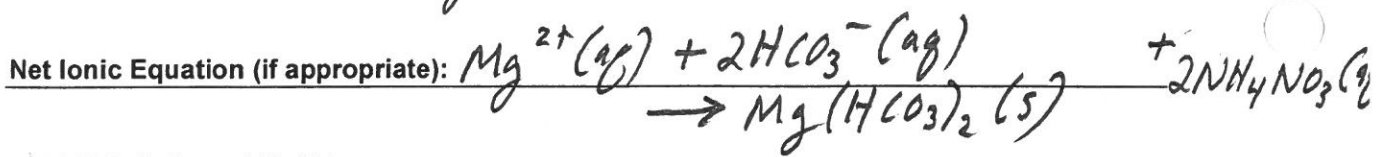
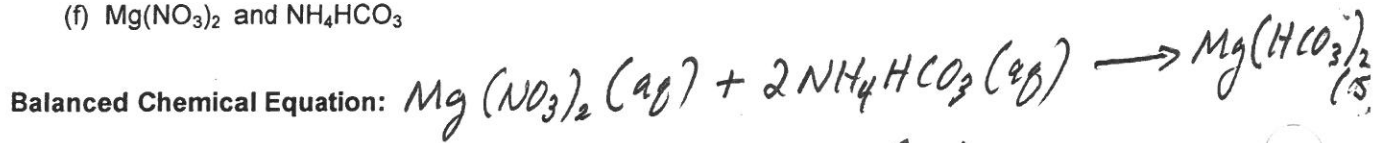
(d) K_2SO_4 and $LiCl$



(e) KOH and $NaNO_3$



(f) $\text{Mg}(\text{NO}_3)_2$ and NH_4HCO_3



(g) $\text{HC}_2\text{H}_3\text{O}_2$ and NaOH

