

WS 4.45

Chemistry: *Ionic Compounds: Formulas and Naming*

Write the name of each of the following compounds.

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|---------------------------------------|-----------|
| 1. NH_4Cl | 1. _____ |
| 2. $\text{NaC}_2\text{H}_3\text{O}_2$ | 2. _____ |
| 3. $\text{Ca}(\text{BrO}_3)_2$ | 3. _____ |
| 4. $\text{Pb}(\text{SO}_4)_2$ | 4. _____ |
| 5. $(\text{NH}_4)_3\text{N}$ | 5. _____ |
| 6. NH_4NO_3 | 6. _____ |
| 7. $\text{Sr}_3(\text{PO}_4)_2$ | 7. _____ |
| 8. BaSO_4 | 8. _____ |
| 9. AgIO_3 | 9. _____ |
| 10. KCN | 10. _____ |
| 11. AuNO_3 | 11. _____ |

Write the chemical formula for each of the given names.

- | | |
|---------------------------|-----------|
| 12. sodium cyanide | 12. _____ |
| 13. barium nitrate | 13. _____ |
| 14. ammonium sulfate | 14. _____ |
| 15. titanium (II) cyanide | 15. _____ |
| 16. calcium phosphate | 16. _____ |
| 17. cesium carbonate | 17. _____ |
| 18. iron (III) acetate | 18. _____ |
| 19. calcium sulfide | 19. _____ |
| 20. beryllium bicarbonate | 20. _____ |
| 21. rubidium iodate | 21. _____ |
| 22. copper (II) fluoride | 22. _____ |
| 23. chromium (VI) oxide | 23. _____ |

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Chemistry: *Ionic Formulas (Binary, Polyatomic, Transition Metals)*

Write the formula for each of the following compounds.

1. sodium iodide 1. _____
2. mercury (II) sulfate 2. _____
3. lead (II) phosphate 3. _____
4. ammonium sulfide 4. _____
5. aluminum chloride 5. _____
6. copper (I) carbonate 6. _____
7. manganese (IV) oxide 7. _____
8. potassium sulfate 8. _____
9. iron (III) oxide 9. _____
10. magnesium nitrate 10. _____
11. calcium sulfide 11. _____
12. potassium cyanide 12. _____
13. magnesium chloride 13. _____
14. chromium (III) oxide 14. _____
15. gold (III) bromide 15. _____
16. beryllium fluoride 16. _____
17. zinc (II) sulfide 17. _____
18. lithium oxide 18. _____
19. aluminum nitrate 19. _____
20. iron (II) phosphide 20. _____
21. magnesium nitrate 21. _____
22. calcium nitrate 22. _____
23. lead (II) chloride 23. _____
24. tin (IV) hydroxide 24. _____
25. sodium oxide 25. _____